

- The metal M forms an oxide M_2O_3 . The formula of its halide will be
a) M_2X_3 b) MX_3 c) M_2X d) M_3X_2
- An element X has electronic configuration 2, 8, 2 and another element Y has electronic configuration 2, 7. They form the compound Z. Formula of Z is
a) X_2Y_3 b) X_3Y_2 c) XY d) XY_2
- Symbol of mercury is
a) Mn b) Mg c) Hg d) Ag
- Latine name of Tin is
a) plumbum b) stannum c) stibium d) wolfram
- Atomicity of ozone is
a) 1 b) 2 c) 3 d) 4
- Molecules of elements are
a) homo atomic b) hetero atomic c) triatomic d) none
- Cation is formed by _____ from neutral atom
a) loss of electron b) addition of proton c) gain of electron d) sharing of electron
- Negatively charged ion is called
a) anion b) cation c) electropositive ion d) b and c
- Formula of sulphuric acid is
a) HNO_3 b) H_2SO_4 c) HCl d) $HClO_4$
- Valency of plumbic ion is
a) 1 b) 2 c) 3 d) 4
- Ammonium ion is
a) NH_4^+ b) Al^{3+} c) N^{3-} d) PH_4^+
- Peroxide ion contains _____ number of oxygen atoms
a) 1 b) 2 c) 3 d) 4
- Formula for aluminium sulphate is
a) $AlSO_4$ b) $Al_2(SO_4)_3$ c) Al_2SO_4 d) $Al(SO_4)_3$
- $5 N_2$ represents
a) 5 nitrogen atoms b) 5 nitrogen molecules c) 10 nitrogen atoms d) 10 nitrogen molecules
- The number of oxygen atom in calcium carbonate is
a) 1 b) 3 c) 5 d) 7
- Valency of sodium is
a) 1 b) 2 c) 3 d) 4
- X has valency 2 and Y has valency 3. What is the formula of compound formed from X and Y
a) X_2Y_3 b) X_3Y_2 c) XY d) none
- Poly atomic molecule is
a) S_8 b) O_3 c) O_2 d) N_2
- Heteroatomic molecule is/are
a) CO_2 b) NH_3 c) P_4 d) a and b
- Chemical formula represents
a) shortest notation of a element or compound
b) the types of elements in a compound
c) number of atoms of each element present in the compound
d) all the above

KEY

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|-----|---|-----|---|-----|---|-----|---|
| 1. | b | 2. | b | 3. | c | 4. | b |
| 5. | c | 6. | a | 7. | a | 8. | a |
| 9. | b | 10. | d | 11. | a | 12. | b |
| 13. | b | 14. | b | 15. | b | 16. | a |
| 17. | b | 18. | a | 19. | d | 20. | d |